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Chapter 7 - Ionic and Metallic Bonding. Jennie L. Borders. Section 7.1 - Ions. Valence electrons are the electrons in the highest occupied energy level. Valence electrons are the only electrons involved in chemical bonding. Elements in the same group have

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the same number of valence electrons.

Chapter 7 - Ionic and Metallic Bonding

Chapter 7 "Ionic and Metallic Bonding" Click to add text 2

Section 7.1 - Ions OBJECTIVES:-Determine the number of valence electrons in an atom of a representative element. 3 Section 7.1 -

Ions OBJECTIVES:-Explain how the octet rule applies to atoms of metallic and nonmetallic elements. 4 Section 7.1 - Ions

OBJECTIVES:-Describe how cations form. 5

Section 7.1 - Ions Chapter 7 "Ionic and Metallic Bonding"

Ionic Bonding. Anions and cations are held together by .

opposite. charges (+ and -) Ionic compounds are called . salts.

Formula unit - Simplest . ratio. of elements in an ionic

compound. Practice with Formula Unit. ... Chapter 7 Ionic and

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Chapter 7 Ionic and Metallic Bonding

Chapter 7: Ionic and Metallic Bonding. STUDY. PLAY. Valence Electrons. electron in the highest occupied energy level of an element's atoms. electron dot structure. diagram that shows valence electrons as dots. octet rule. in a chemical reaction, atoms gain or lose electrons to acquire the electron structure of a noble gas.

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Chapter 7 - Ionic and Metallic Bonding. STUDY. Flashcards. Learn. Write. Spell. Test. PLAY. Match. Gravity. Created by. ... b.It is held together by ionic bonds, c.It is composed of anions and

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cations. d.All of the above. Solid. ... What is the basis of a metallic bond?

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Chemistry, Chapter 7, Ionic & Metallic Bonding, Review DRAFT. 9th - 12th grade. 68 times. Chemistry. 52% average accuracy. 3 years ago. kirch. 0. Save. Edit. ... Which of the following pairs of elements will form an ionic bond? answer choices . K and Ca. Co and Ni. F and S. Sr and Br. Tags: Question 33 . SURVEY . 30 seconds . Q. Which IS a ...

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Chemistry, Chapter 7, Ionic & Metallic Bonding, Review ...

(a) Ionic Bonding: Atoms are held together by electrostatic forces between two oppositely charged ions. Covalent Bonding: Atoms share electrons (in valence shell) with each other so as to attain a stable electronic configuration. Metallic Bonding: Positively charged metallic ions are held together surrounded by a cloud of electrons.

(a) Briefly cite the main differences between ionic ...

Metallic Bonds and Metallic Properties

1. Is the following sentence true or false? Metals are made up of cations and valence electrons, not neutral atoms.
2. What are metallic bonds?
3. Name three properties of metals that can be explained by metallic bonding. a. b. c.
4. What happens to an ionic crystal when a force is applied to it?
- 5.

BONDING AND INTERACTIONS

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IONIC AND METALLIC BONDING. Chapter Test A. A. Matching. Match each description in Column B with the correct term in Column A. Write the letter of the correct description on the line. Column A. E 1. electron dot structure. F 2. ionic compound. H 3. valence electron. C 4. ionic bond. J 5. chemical formula. G 6. halide ion.

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Chapter 7 - Ionic and Metallic Bonding - 7 Assessment - Page 214: 46 Answer Noble gases are excluded from Table 6.2 (Electronegativity values) because they do not form many compounds, or in other words, they rarely react.

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Ch 7 Ionic And Metallic Bonding Answers

Chapter 10: Chemical Bonding I: Basic Concepts Compounds described in Example 10-2 are binary ionic compounds: consists of monoatomic cations and monoatomic anions. Ternary ionic compounds consists of monoatomic and polyatomic ions. Bonding between atoms within the polyatomic ions is covalent. With the exception of ion pairs, such as $(- \dot{\text{Cl}} + \dot{\text{Na}} \dot{\text{Cl}})$, that may be found in the gaseous ...

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